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**1 Mark Each**

- Heat is evolved during
  - Combination Reaction
  - Combustion Reaction
  - Endothermic Reaction
  - Displacement Reaction
- Dissolving sugar in water is an example of
  - Physical Change
  - Redox Reaction
  - Chemical change
  - None of the given three
- Chemically rust is
  - Hydrated ferric oxide
  - Hydrated ferrous oxide
  - only ferrous oxide
  - only ferric oxide
- The reaction  $\text{H}_2 + \text{Cl}_2 \longrightarrow 2\text{HCl}$  represents
  - Combustion
  - Combination
  - Reduction
  - Oxidation
- Oxidation is a process which involves :
  - Addition of oxygen
  - Removal of hydrogen
  - Loss of electrons
  - All are correct

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6. Which of the following is not a balanced equation?
- (a)  $\text{H}_2 + \text{O}_2 \longrightarrow \text{H}_2\text{O}$   
(b)  $\text{Mg} + \text{CuSO}_4 \longrightarrow \text{MgSO}_4 + \text{Cu}$   
(c)  $\text{NaOH} + \text{HCl} \longrightarrow \text{NaCl} + \text{H}_2\text{O}$   
(d)  $\text{Zn} + \text{S} \longrightarrow \text{ZnS}$ .
7. Copper displacement which of the following metals from its salt solution:
- (a)  $\text{ZnSO}_4$   
(b)  $\text{FeSO}_4$   
(c)  $\text{AgNO}_3$   
(d)  $\text{NiSO}_4$
8. PbS reacts with ozone and forms  $\text{PbSO}_4$ . As per the balanced equation, molecules of ozone required for every one molecule of PbS are
- (a) 6  
(b) 4  
(c) 3  
(d) 2
9. Some crystals of copper sulphate were dissolved in water. The colour of the solution obtained would be:
- (a) green  
(b) red  
(c) blue  
(d) brown
10. When dilute hydrochloric acid is added to zinc pieces taken in a test tube:
- (a) no change takes place  
(b) the colour of the solution becomes yellow  
(c) a pungent smelling gas gets liberated  
(d) small bubbles of hydrogen gas appear on the surface of zinc pieces.

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